

Quiz 1 topic Reminder

I understand the value of K and can generate an equilibrium expression

(#11-1) I understand the value of K and can generate an equilibrium expression

(#11-1a)

- I can write an equilibrium constant expression K_c or K_{eq}
- I can write an equilibrium constant expression K_p (partial pressures.)
- I can write an equilibrium constant expression for solubility K_{sp} .

(#11-1b)

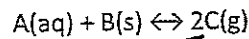
- I can explain how the value of the equilibrium constant (k) relates to the extent of the reaction.

(#11-1c)

- I can calculate a new K value from an old K value if the reaction is altered (Altered means reversed or coefficients are a multiple of the original.)

(#11-1d)

- I can specifically model an equilibrium system using particulate diagrams.
- I can determine if a reaction is at equilibrium.



1. Write the equilibrium expression

a. $K_c = \frac{[C]^2}{[A]}$

b. $K_p = \rightarrow K_p = (P_c)^2$

2. In a 2L rigid vessel, 0.2M A and solid B are added. Since the reaction is slightly product favored, draw a graph below and rigid container representing the reaction at equilibrium.

3.

