Quiz 1 topic Reminder

I understand the value of K and can generate an equilibrium expression

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(#11-1) I understand the value of K and can generate an equilibrium expression

(#11-1a)

- I can write an equilibrium constant expression Kc or Keq
- I can write an equilibrium constant expression Kp (partial pressures.)
- I can write an equilibrium constant expression for solubility Ksp.

(#11-1b)

- I can explain how the value of the equilibrium constant (k) relates to the extent of the reaction.

(#11-1c)

- I can calculate a new K value from an old K value if the reaction is altered (Altered means reversed or coefficients are a multiple of the original.)

(#11-1d)

- I can specifically model an equilibrium system using particulate diagrams.
- I can determine if a reaction is at equilibrium.

$$A(aq) + B(s) \leftrightarrow 2C(g)$$

1. Write the equilibrium expression
   a. Kc = $\frac{[C]^2}{[A]}$
   b. Kp = $\longrightarrow K_p = (P_C)^2$
2. In a 2L rigged vessel, 0.2M A and solid B are added. Since the reaction is slightly product favored, draw a graph below and rigged container representing the reaction at equilibrium.
3.